KEY

Write the correct formula for the compound formed by each of the following ion pairs.

1. NH_4^+ and S^{2-} $(NH_4)_2S$
2. Al^{3+} and CrO_4^{2-} $Al_2(CrO_4)_3$
3. Mg^{2+} and ClO_3^- $Mg(ClO_3)_2$
4. Ca^{2+} and PO_4^{3-} $Ca_3(PO_4)_2$

Name the following compounds.

5. Na_2S Sodium Sulfide
6. NH_4NO_2 Ammonium Nitrite
7. $MgBr_2$ Magnesium Bromide
8. $PbSO_4$ Lead(II) Sulfate
9. CuF_2 Copper(II) fluoride
10. N_2O_4 dinitrogen tetroxide
11. P_2O_5 diphosphorous pentoxide

Write balanced formulas for each of the following compounds.

12. Copper (II) chloride Cu^{2+} Cl^- $CuCl_2$
13. Ammonium sulfate NH_4^+ SO_4^{2-} $(NH_4)_2SO_4$
14. trinitrogen monoxide N_3O
15. diphosphorus pentoxide P_2O_5
16. magnesium chloride Mg^{2+} O^{2-} MgO
17. lead (IV) oxide Pb^{4+} O^{2-} $Pb_2O_4 \Rightarrow PbO_2$
18. potassium arsenate K^+ AsO_3^{3-} K_3AsO_3
19. calcium cyanide Ca^{2+} CN^- $Ca(CN)_2$
20. hexachlorine dibromide Cl_6Br_2