

Types of Reactions Lab - Analysis Key

Station	Balanced Chemical Equation	Type of Reaction
Demmo 1	$Mg + O_2 \rightarrow MgO$	Synthesis
Demmo 2	$2H_2 + O_2 \rightarrow 2H_2O$	Synthesis
1.1	$Mg + KOH \rightarrow N.R.$	SR
1.2	$3Mg + Al_2(SO_4)_3 \rightarrow 3MgSO_4 + 2Al$	SR
1.3	$Mg + CuSO_4 \rightarrow MgSO_4 + Cu$	SR
1.4	$Mg + NiCl_2 \rightarrow MgCl_2 + Ni$	SR
1.5	$Al + KOH \rightarrow N.R.$	SR
1.6	$Al + Al_2(SO_4)_3 \rightarrow N.R.$	SR
1.7	$2Al + 3CuSO_4 \rightarrow Al_2(SO_4)_3 + 3Cu$	SR
1.8	$2Al + 3NiCl_2 \rightarrow 2AlCl_3 + 3Ni$	SR
1.9	$Cu + KOH \rightarrow N.R.$	SR
1.10	$Cu + Al_2(SO_4)_3 \rightarrow N.R.$	SR
1.11	$Cu + CuSO_4 \rightarrow N.R.$	SR
1.12	$2Cu + NiCl_2 \rightarrow 2CuCl + Ni$ or $Cu + NiCl_2 \rightarrow CuCl_2 + Ni$	SR

