STEM Chem Unit 2 Test part 2 REVIEW Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ per\_\_\_\_\_\_\_\_\_\_\_\_\_

|  |  |
| --- | --- |
|  | IS: I can evaluate experimental methods and comment of sources of error. |
|  | AST 9.6: I can balance chemical equations to uphold the Law of Conservation of Mass |
|  | AST 9.7: I can predict the theoretical yield of a chemical reaction and compare it to the experimental yield (determining % yield) |
|  | ALT 10: I can apply stoichiometry in the calculation of reactant and product quantities. |
|  | AST 10.6: I can identify limiting and excess reagents in a chemical reaction. |

**Silver Lab Data:**

|  |  |
| --- | --- |
| Part 1 |  |
| Mass of empty beaker |  |
| Mass of original silver nitrate |  |
| Mass of beaker with silver |  |
| Part 2 |  |
| Mass of beaker after treating with excess nitric acid (HNO3) |  |
| Mass of sodium chloride used |  |
| Mass of filter paper |  |
| Mass of filter paper with final product |  |

**Silver Lab basic calculations:**

|  |  |
| --- | --- |
| Mass of silver produced |  |
| Mass of product after treating with excess HNO3 |  |
| Mass of final product produced |  |

REVIEW:

You will need to be able to determine the percent yield for each of the three reactions that we did in this lab.

You will need to be able to determine the percent error for the entire lab (this means you will need to determine the theoretical yield of your final product starting from the original silver nitrate)

You will also need to be able to do an error analysis for the lab.

\*\*\*You will be able to use this data sheet on the test but nothing else, so do some good reviewing. You may record the skeleton equations for the three reactions here. No other notes on this page will be permitted.

Rxn A:

Rxn B:

Rxn. C: