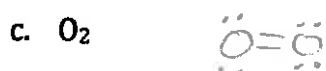
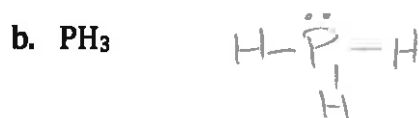


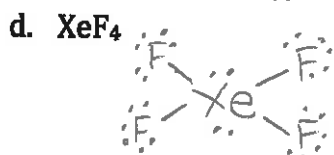
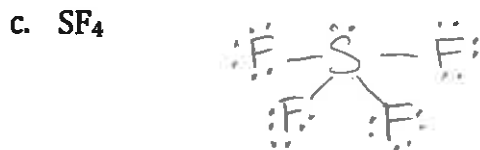
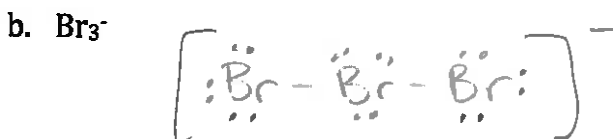
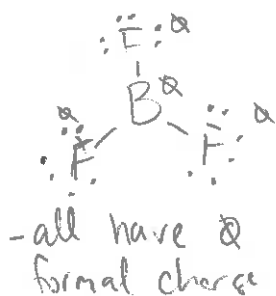
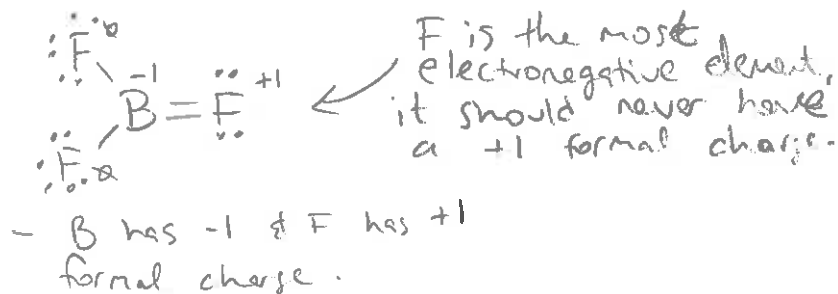
STEM Lewis Dot Structure Practice

Name KEY

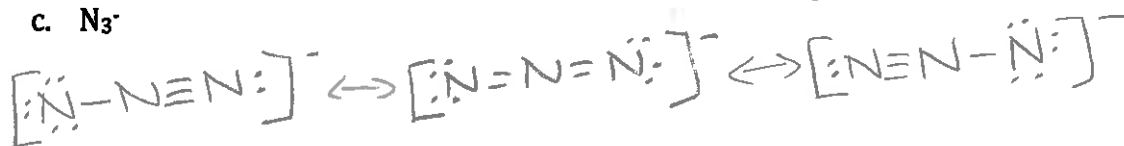
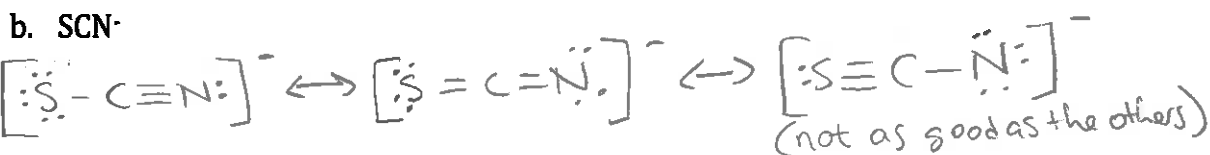
1) Write the Lewis structures for the following molecules (the central atom obeys the octet rule).



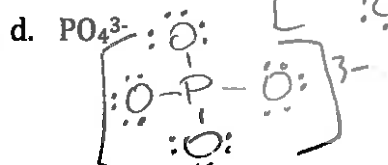
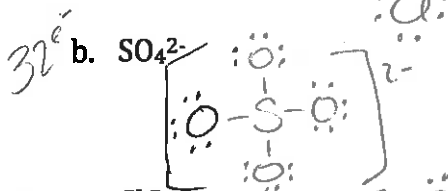
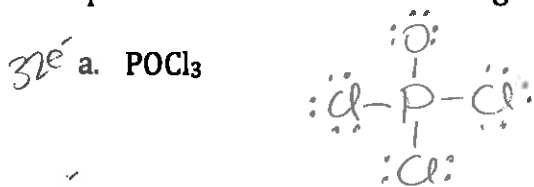
2) Write Lewis Structures for the following molecules (the central atom does not obey the octet rule).

3) Use the formal charge arguments to rationalize why BF_3 would not follow the octet rule.24 e⁻

4) Write Lewis structures for the following. Show all resonance structures where applicable.



5) Write the Lewis structures that obey the octet rule for the following compounds. Show the formal charge for each central atom.



6) Write the best Lewis structures for the compounds in number 5 (involve minimum formal charges, may not obey the octet rule)

